

Set # 4: Write formulas for the following mixed

I = ionic
C = covalent
A = acid

compounds $\text{Co}^{+2}\text{ClO}_4^{-1}$ $\text{Co}(\text{ClO}_4)_2$

I1.) Cobalt (II) perchlorate ~~$\text{Co}^{+2}\text{ClO}_4^{-1}$~~ ~~$\text{Co}(\text{ClO}_4)_2$~~

C2.) Tetracarbon Heptasulfide C_4S_7

I3.) Sodium chlorite $\text{Na}^{+1}\text{ClO}_2^{-1}$ NaClO_2

A4.) Hydroiodic acid HI

I5.) Strontium bromide $\text{Sr}^{+2}\text{Br}^{-1}$ SrBr_2

C6.) Sulfur Nonoxide SO_9

I7.) Cadmium oxide $\text{Cd}^{+2}\text{O}^{-2}$ CdO

C8.) Silicon tetrabromide SiBr_4

I9.) Sodium bicarbonate $\text{Na}^{+1}\text{HCO}_3^{-1}$ $= \text{NaHCO}_3$

10.) Tin (IV) nitrate $\text{Sn}^{+4}\text{NO}_3^{-1}$ $\text{Sn}(\text{NO}_3)_4$