

Set # 2: Write formulas for the following mixed

C = covalent
I = ionic
A = acid

compounds

C 1.) Phosphorus pentachloride PCl_5

I 2.) Aluminum nitrite $\text{Al}^{+3} \text{NO}_2^{-1}$ ~~$\text{Al}^{+3} \text{NO}_2^{-1}$~~ $\text{Al}(\text{NO}_2)_3$

A 3.) Nitrous acid $\text{H}^{+1} \text{NO}_2^{-1}$ HNO_2

I 4.) Tin (II) sulfate ~~$\text{Sn}^{+2} \text{SO}_4^{-2}$~~ SnSO_4

I 5.) Tin (IV) sulfate ~~$\text{Sn}^{+4} \text{SO}_4^{-2}$~~ $\text{Sn}(\text{SO}_4)_2$

C 6.) Trisulfur hexoxide S_3O_6

A 7.) Carbonic acid $\text{H}_2^{+1} \text{CO}_3^{-2}$ H_2CO_3

I 8.) Magnesium sulfide ~~$\text{Mg}^{+2} \text{S}^{-2}$~~ MgS

A 9.) Acetic acid $\text{H}^{+1} \text{C}_2\text{H}_3\text{O}_2^{-1}$ $\text{HC}_2\text{H}_3\text{O}_2$

I 10.) Potassium phosphide $\text{K}_3^{+1} \text{P}^{-3}$ K_3P