

Set # 1: Name the following mixed compounds.

1.)  $\text{BaCrO}_4$  barium chromate      6.)  $\text{HClO}_4$  perchloric acid

2.)  $\text{HF}$  hydrofluoric acid      7.)  $\text{Na}_2\text{O}$  sodium oxide

3.)  $\text{N}_2\text{O}_4$  dinitrogen tetroxide      8.)  $\text{CoCl}_2$  cobalt (II) chloride

4.)  $\text{CCl}_4$  carbon tetrachloride      9.)  $\text{Ca(OH)}_2$  calcium hydroxide

5.)  $\text{CuO}$  copper (II) oxide      10.)  $\text{SiO}_2$  silicon dioxide

Set # 2: Write formulas for the following mixed

compounds

1. ) Phosphorus pentachloride  $\text{PCl}_5$
2. ) Aluminum nitrite  $\text{Al}(\text{NO}_2)_3$
3. ) Nitrous acid  $\text{HNO}_2$
4. ) Tin (II) sulfate  $\text{SnSO}_4$
5. ) Tin (IV) sulfate  $\text{Sn}(\text{SO}_4)_2$
6. ) Trisulfur hexoxide  $\text{S}_3\text{O}_6$
7. ) Carbonic acid  $\text{H}_2\text{CO}_3$
8. ) Magnesium sulfide  $\text{MgS}$
9. ) Acetic acid  $\text{HC}_2\text{H}_3\text{O}_2$  or  $\text{HCH}_3\text{COO}$
- 10.) Potassium phosphide  $\text{K}_3\text{P}$