

**Naming Worksheet #3 - solutions**

Name the following chemical compounds:

- |     |                                                                |                                    |
|-----|----------------------------------------------------------------|------------------------------------|
| 1)  | RbBr                                                           | <b>rubidium bromide</b>            |
| 2)  | Ca <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub>                | <b>calcium phosphate</b>           |
| 3)  | NiF <sub>3</sub>                                               | <b>nickle (III) fluoride</b>       |
| 4)  | Sn(NO <sub>3</sub> ) <sub>4</sub>                              | <b>tin (IV) nitrate</b>            |
| 5)  | Fe(C <sub>2</sub> H <sub>3</sub> O <sub>2</sub> ) <sub>3</sub> | <b>iron (III) acetate</b>          |
| 6)  | K <sub>3</sub> P                                               | <b>potassium phosphide</b>         |
| 7)  | N <sub>2</sub> O <sub>3</sub>                                  | <b>dinitrogen trioxide</b>         |
| 8)  | Cu(NO <sub>2</sub> ) <sub>2</sub>                              | <b>copper (II) nitrite</b>         |
| 9)  | P <sub>4</sub> Se <sub>3</sub>                                 | <b>tetraphosphorus triselenide</b> |
| 10) | (NH <sub>4</sub> ) <sub>2</sub> SO <sub>4</sub>                | <b>ammonium sulfate</b>            |

Write the formulas for the following chemical compounds:

- |     |                                |                                                   |
|-----|--------------------------------|---------------------------------------------------|
| 11) | lithium bromide                | <b>LiBr</b>                                       |
| 12) | beryllium nitride              | <b>Be<sub>3</sub>N<sub>2</sub></b>                |
| 13) | carbon tetrachloride           | <b>CCl<sub>4</sub></b>                            |
| 14) | dihydrogen monoxide            | <b>H<sub>2</sub>O</b>                             |
| 15) | copper (I) phosphate           | <b>Cu<sub>3</sub>PO<sub>4</sub></b>               |
| 16) | magnesium sulfate heptahydrate | <b>MgSO<sub>4</sub>·7H<sub>2</sub>O</b>           |
| 17) | lead (II) phosphate            | <b>Pb<sub>3</sub>(PO<sub>4</sub>)<sub>2</sub></b> |
| 18) | diselenium diiodide            | <b>Se<sub>2</sub>I<sub>2</sub></b>                |
| 19) | disilicon hexabromide          | <b>Si<sub>2</sub>Br<sub>6</sub></b>               |
| 20) | copper (II) phosphide          | <b>Cu<sub>3</sub>P<sub>2</sub></b>                |

## Naming Acids and Bases Answers

Name the following acids and bases:

- 1) NaOH      sodium hydroxide
- 2) H<sub>2</sub>SO<sub>3</sub>      sulfurous acid
- 3) H<sub>2</sub>S      hydrosulfuric acid
- 4) H<sub>3</sub>PO<sub>4</sub>      phosphoric acid
- 5) Ca(OH)<sub>2</sub>      calcium hydroxide
- 6) Fe(OH)<sub>3</sub>      iron (III) hydroxide
- 7) H<sub>3</sub>P      hydrophosphoric acid

Write the formulas of the following acids and bases:

- 8) hydrofluoric acid      HF
- 9) hydroselenic acid      H<sub>2</sub>Se
- 10) carbonic acid      H<sub>2</sub>CO<sub>3</sub>
- 11) lithium hydroxide      LiOH
- 12) nitrous acid      HNO<sub>2</sub>
- 13) cobalt hydroxide      Co(OH)<sub>2</sub>
- 14) sulfuric acid      H<sub>2</sub>SO<sub>4</sub>
- 15) beryllium hydroxide      Be(OH)<sub>2</sub>
- 16) hydrobromic acid      HBr